



Rewarding Learning

ADVANCED
General Certificate of Education
2016

Centre Number

--	--	--	--	--

Candidate Number

--	--	--	--	--

Chemistry

Assessment Unit A2 3

assessing

Module 3:

Practical Examination

Practical Booklet B

[AC234]

AC234

FRIDAY 13 MAY, MORNING

TIME

1 hour 15 minutes.

INSTRUCTIONS TO CANDIDATES

Write your Centre Number and Candidate Number in the spaces provided at the top of this page.

You must answer the questions in the spaces provided.

Do not write outside the boxed area on each page or on blank pages.

Complete in blue or black ink only. **Do not write with a gel pen.**

Answer **all three** questions.

INFORMATION FOR CANDIDATES

The total mark for this paper is 50.

Question 1 is a practical exercise worth 17 marks.

Question 2 is a practical exercise worth 13 marks.

Question 3 is a planning exercise worth 20 marks.

Quality of written communication will be assessed in **Question 3(e)**.

Figures in brackets printed down the right-hand side of pages indicate the marks awarded to each question or part question.

A Periodic Table of Elements (including some data) is provided.

10138



12AC23401

(c) (i) Write the half-equation for the reduction of iodate(V) ions, in the presence of hydrogen ions, to form iodine.

_____ [2]

(ii) Write the overall ionic equation for the reaction of iodate(V) ions with iodide ions, in the presence of hydrogen ions, to give iodine and water.

_____ [2]

(d) Calculate the concentration of the potassium iodate(V) solution in g dm^{-3} .

_____ [4]

[Turn over



2 (a) (i) Based on the following observations, make deductions for salts **X** and **Y**.

Test	Observation	Deductions
<p>1 Dissolve two spatula measures of X in 50 cm³ of water.</p> <p>Keep this solution for use in further tests.</p>	<p><i>Green crystals dissolve to produce a green solution.</i></p>	[1]
<p>2 Dissolve two spatula measures of Y in 50 cm³ of water.</p> <p>Keep this solution for use in further tests.</p>	<p><i>Pink crystals dissolve to produce a pink solution.</i></p>	
<p>3 Place 4 cm³ of the solution of X in a test tube. Slowly add an equal volume of sodium hydroxide solution.</p> <p>Add concentrated ammonia.</p>	<p><i>A green precipitate forms.</i></p> <p><i>The precipitate disappears and a blue solution is formed.</i></p>	[1]
<p>4 Place 4 cm³ of the solution of Y in a test tube. Slowly add concentrated ammonia until present in excess.</p> <p>Shake the solution.</p>	<p><i>A blue precipitate forms. The precipitate disappears and a yellow solution is formed.</i></p> <p><i>The yellow solution turns brown.</i></p>	[1]
<p>5 Place 4 cm³ of the solution Y in a test tube. Add 2 cm³ of concentrated hydrochloric acid.</p> <p>Add 8 cm³ of water and shake vigorously.</p>	<p><i>Solution turns blue and back to pink.</i></p>	
<p>6 Slowly add a solution of 1,2-diaminoethane to the solution of X until it is present in excess.</p>	<p><i>Solution turns purple.</i></p>	



(ii) Give the formula of the green precipitate in Test 3.

_____ [1]

(iii) Give the formula of the species responsible for the blue solution in Test 3.

_____ [1]

(iv) Give the formula of the species responsible for the yellow solution in Test 4.

_____ [1]

(v) Give the equation for the reaction in Test 5.

_____ [2]

(vi) Give the formula of the species responsible for the purple colour in Test 6.

_____ [1]

[Turn over



(b) (i) Based on the following observations, make deductions for compound **Z**.

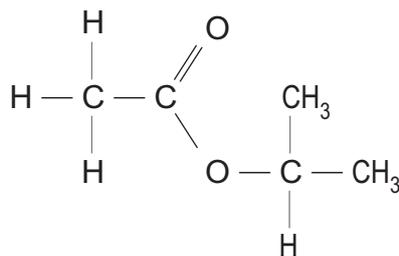
Test	Observation	Deductions
1 Add 2 cm ³ of Z to 2 cm ³ of water.	<i>One layer formed.</i>	
2 Add sodium hydrogencarbonate to Z .	<i>No bubbles produced.</i>	[1]
3 Add a few drops of Z to a solution of 2,4-dinitrophenylhydrazine.	<i>An orange solid forms.</i>	[1]
4 Heat Z with Tollen's reagent.	<i>The solution remains colourless.</i>	[1]

(ii) Suggest a possible structure for **Z**.

[1]



3 Isopropyl acetate is an ester with a boiling point of 88 °C.



isopropyl acetate

It can be prepared by refluxing acetic acid with isopropyl alcohol in the presence of a catalyst in a round-bottomed flask. The product is removed from the flask, purified and analysed.

(a) (i) Write an equation for the *equilibrium* reaction between acetic acid and isopropyl alcohol.

_____ [2]

(ii) Name the catalyst added to the flask.

_____ [1]

(iii) What else should be added to the round-bottomed flask?

_____ [1]

(b) Give the IUPAC name for isopropyl alcohol.

_____ [1]

[Turn over



(c) Assuming a 40% yield, what is the minimum mass of isopropyl alcohol required to produce 10.2g of isopropyl acetate?

[4]

(d) Name the experimental technique used to remove the product from the flask.

[1]

(e) Giving experimental details, describe how the crude product can be purified

(i) Using an aqueous solution of sodium carbonate

[3]



(ii) Using anhydrous calcium chloride.

[3]

Quality of written communication

[2]

(f) Suggest how infrared spectroscopy could be used to show that the product did not contain any unreacted acetic acid or isopropyl alcohol.

[1]

(g) State the integration pattern in the nmr spectrum of isopropyl acetate.

[1]

THIS IS THE END OF THE QUESTION PAPER



BLANK PAGE
DO NOT WRITE ON THIS PAGE

10138



12AC23410





BLANK PAGE
DO NOT WRITE ON THIS PAGE

10138



12AC23411

DO NOT WRITE ON THIS PAGE

For Examiner's use only		
Question Number	Examiner Mark	Remark
1		
2		
3		
Total Marks		

Permission to reproduce all copyright material has been applied for.
In some cases, efforts to contact copyright holders may have been unsuccessful and CCEA
will be happy to rectify any omissions of acknowledgement in future if notified.

212738



12AC23412